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5. Converting from Moles to Mass Using the amounts of silicon you calculated above, determine the amount in moles of carbon monoxide, CO, produced. Then convert the amounts of CO in moles into masses in grams. You can use the equations shown below. Substitute your numbers for the x 's. Copy and paste these set-up equations if needed to show the work for each of your calculations.

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The Bunsen burner is the most common source of heat in the chemistry laboratory. Natural gas enters the burner near the base, rises through a barrel, mixes with air, and burns.

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